

ENAJČBE KEMIJSKIH REAKCIJ (Zapišite ionsko obliko enačb reakcij, ki so potekle, v ostalih primerih samo naštejete ione v posamezni kapljici. V tabeli sta navedena primera.)

Št.	Ionska oblika enačbe reakcije ali naštejete ione
1.	$\text{H}_3\text{O}^+(\text{aq}) + \text{Cl}^-(\text{aq}) + \text{NH}_4^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow 2 \text{H}_2\text{O}(\text{l}) + \text{Cl}^-(\text{aq}) + \text{NH}_4^+(\text{aq})$
2.	$\text{H}_3\text{O}^+(\text{aq}), \text{Cl}^-(\text{aq}), \text{Ca}^{2+}(\text{aq})$
3.	$\text{H}_3\text{O}^+(\text{aq}), \text{Cl}^-(\text{aq}), \text{Cu}^{2+}(\text{aq}), \text{SO}_4^{2-}(\text{aq})$
4.	$\text{H}_3\text{O}^+(\text{aq}) + \text{Cl}^-(\text{aq}) + \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow 2 \text{H}_2\text{O}(\text{l}) + \text{Cl}^-(\text{aq}) + \text{Na}^+(\text{aq})$
5.	$\text{H}_3\text{O}^+(\text{aq}), \text{Cl}^-(\text{aq}), \text{Ba}^{2+}(\text{aq})$
6.	$2\text{H}_3\text{O}^+(\text{aq}) + 2\text{Cl}^-(\text{aq}) + 2\text{Na}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow 3 \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g}) + 2\text{Cl}^-(\text{aq}) + 2\text{Na}^+(\text{aq})$
7.	$\text{H}_3\text{O}^+(\text{aq}) + \text{Cl}^-(\text{aq}) + \text{Ag}^+(\text{aq}) + \text{NO}_3^-(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{H}_3\text{O}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$
8.	$2\text{OH}^-(\text{aq}) + 2\text{NH}_4^+(\text{aq}) + \text{Ca}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq}) \rightarrow \text{Ca}(\text{OH})_2(\text{s}) + 2\text{NH}_4^+(\text{aq}) + 2\text{Cl}^-(\text{aq})$
9.	$2\text{OH}^-(\text{aq}) + 2\text{NH}_4^+(\text{aq}) + \text{Cu}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) \rightarrow \text{Cu}(\text{OH})_2(\text{s}) + 2\text{NH}_4^+(\text{aq}) + \text{SO}_4^{2-}(\text{aq})$
10.	$\text{NH}_4^+(\text{aq}), \text{OH}^-(\text{aq}), \text{Na}^+(\text{aq})$
11.	$2\text{OH}^-(\text{aq}) + 2\text{NH}_4^+(\text{aq}) + \text{Ba}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq}) \rightarrow \text{Ba}(\text{OH})_2(\text{s}) + 2\text{NH}_4^+(\text{aq}) + 2\text{Cl}^-(\text{aq})$
12.	$\text{NH}_4^+(\text{aq}), \text{OH}^-(\text{aq}), \text{Na}^+(\text{aq}), \text{CO}_3^{2-}(\text{aq})$
13.	$\text{NH}_4^+(\text{aq}) + \text{OH}^-(\text{aq}) + \text{Ag}^+(\text{aq}) + \text{NO}_3^-(\text{aq}) \rightarrow \text{AgOH}(\text{s}) + \text{NH}_4^+(\text{aq}) + \text{NO}_3^-(\text{aq})$
14.	$\text{Ca}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq}) + \text{Cu}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) \rightarrow \text{CaSO}_4(\text{s}) + 2\text{Cl}^-(\text{aq}) + \text{Cu}^{2+}(\text{aq})$
15.	$\text{Ca}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq}) + 2\text{Na}^+(\text{aq}) + 2\text{OH}^-(\text{aq}) \rightarrow \text{Ca}(\text{OH})_2(\text{s}) + 2\text{Na}^+(\text{aq}) + 2\text{Cl}^-(\text{aq})$
16.	$\text{Ca}^{2+}(\text{aq}), \text{Cl}^-(\text{aq}), \text{Ba}^{2+}(\text{aq})$
17.	$\text{Ca}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq}) + 2\text{Na}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{CaCO}_3(\text{s}) + 2\text{Na}^+(\text{aq}) + 2\text{Cl}^-(\text{aq})$
18.	$\text{Ca}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq}) + 2\text{Ag}^+(\text{aq}) + 2\text{NO}_3^-(\text{aq}) \rightarrow 2\text{AgCl}(\text{s}) + \text{Ca}^{2+}(\text{aq}) + 2\text{NO}_3^-(\text{aq})$
19.	$\text{Cu}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + 2\text{Na}^+(\text{aq}) + 2\text{OH}^-(\text{aq}) \rightarrow \text{Cu}(\text{OH})_2(\text{s}) + 2\text{Na}^+(\text{aq}) + \text{SO}_4^{2-}(\text{aq})$
20.	$\text{Cu}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + \text{Ba}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq}) \rightarrow \text{BaSO}_4(\text{s}) + \text{Cu}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq})$
21.	$\text{Cu}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + 2\text{Na}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{CuCO}_3(\text{s}) + 2\text{Na}^+(\text{aq}) + \text{SO}_4^{2-}(\text{aq})$
22.	$\text{Cu}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + 2\text{Ag}^+(\text{aq}) + 2\text{NO}_3^-(\text{aq}) \rightarrow \text{Ag}_2\text{SO}_4(\text{s}) + \text{Cu}^{2+}(\text{aq}) + 2\text{NO}_3^-(\text{aq})$
23.	$2\text{Na}^+(\text{aq}) + 2\text{OH}^-(\text{aq}) + \text{Ba}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq}) \rightarrow \text{Ba}(\text{OH})_2(\text{s}) + 2\text{Na}^+(\text{aq}) + 2\text{Cl}^-(\text{aq})$
24.	$\text{Na}^+(\text{aq}), \text{OH}^-(\text{aq}), \text{CO}_3^{2-}(\text{aq})$
25.	$\text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq}) + \text{Ag}^+(\text{aq}) + \text{NO}_3^-(\text{aq}) \rightarrow \text{AgOH}(\text{s}) + \text{Na}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$
26.	$\text{Ba}^{2+}(\text{aq}) + 2\text{Na}^+(\text{aq}) + 2\text{Cl}^-(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{BaCO}_3(\text{s}) + 2\text{Na}^+(\text{aq}) + 2\text{Cl}^-(\text{aq})$
27.	$\text{Ba}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq}) + 2\text{Ag}^+(\text{aq}) + 2\text{NO}_3^-(\text{aq}) \rightarrow 2\text{AgCl}(\text{s}) + \text{Ba}^{2+}(\text{aq}) + 2\text{NO}_3^-(\text{aq})$
28.	$2\text{Na}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) + 2\text{Ag}^+(\text{aq}) + 2\text{NO}_3^-(\text{aq}) \rightarrow \text{Ag}_2\text{CO}_3(\text{s}) + 2\text{Na}^+(\text{aq}) + 2\text{NO}_3^-(\text{aq})$

OPAŽANJA IN SKLEPI

	NH ₃	CaCl ₂	CuSO ₄	NaOH	BaCl ₂	Na ₂ CO ₃	AgNO ₃
HCl	1 H ₂ O	2 /	3 /	4 H ₂ O	5 /	6 CO ₂	7 bela oborina AgCl
	NH ₃	8 bela oborina Ca(OH) ₂	9 modra oborina* Cu(OH) ₂	10 /	11 bela oborina Ba(OH) ₂	12 /	13 rumenkasta oborina AgOH
		CaCl ₂	14 bela oborina CaSO ₄	15 bela oborina Ca(OH) ₂	16 /	17 bela oborina CaCO ₃	18 bela oborina AgCl
			CuSO ₄	19 modra oborina* Cu(OH) ₂	20 bela oborina BaSO ₄	21 modra oborina CuCO ₃	22 bela oborina Ag ₂ SO ₄
				NaOH	23 bela oborina Ba(OH) ₂	24 /	25 rumenkasta oborina AgOH
					BaCl ₂	26 bela oborina BaCO ₃	27 bela oborina AgCl
						Na ₂ CO ₃	28 rumena oborina Ag ₂ CO ₃
							AgNO ₃

ali če je oborina
temna nastane
koordinacijska spojina
tetraamini bakrov(2+) ion

